Some Common Ions

Cations		Anions	
ammonium	NH_4^+	acetate	$C_2H_3O_2^-$
copper (I)	Cu^+	chlorate	ClO_3^-
silver	Ag^+	chlorite	ClO_2^-
cadmium	Cd^{2+}	cyanide	CN^-
chromium (II)	Cr^{2+}	hydroxide	OH^-
cobalt (II)	Co^{2+}	hypochlorite	ClO^-
copper (II)	Cu^{2+}	nitrate	NO_3^-
iron (II)	Fe^{2+}	nitrite	NO_2^-
lead (II)	Pb^{2+}	perchlorate	ClO_4^-
manganese (II)	Mn^{2+}	permanganate	MnO_4^-
mercury (I)	Hg_2^{2+}	carbonate	CO_3^{2-}
mercury (II)	Hg^{2+}	chromate	CrO_4^{2-}
nickel (II)	Ni^{2+}	dichromate	$Cr_2O_7^{2-}$
tin (II)	Sn^{2+}	peroxide	O_2^{2-}
zinc	Zn^{2+}	sulfate	SO_4^{2-}
aluminum	Al^{3+}	sulfite	SO_3^{2-}
chromium (III)	Cr^{3+}	phosphate	PO_4^{3-}
iron(III)	Fe^{3+}	phosphite	PO_3^{3-}

Certain polyatomic anions having charges of "2—" or "3—" may combine with one or more H^+ ions. When only one hydrogen is present, the ion is most simply named by using the prefix bi or just write hydrogen before the ion's regular name. If two hydrogens are present, dihydrogen is used as a prefix.